



Charging a cell phone with a glass of fruit juice?

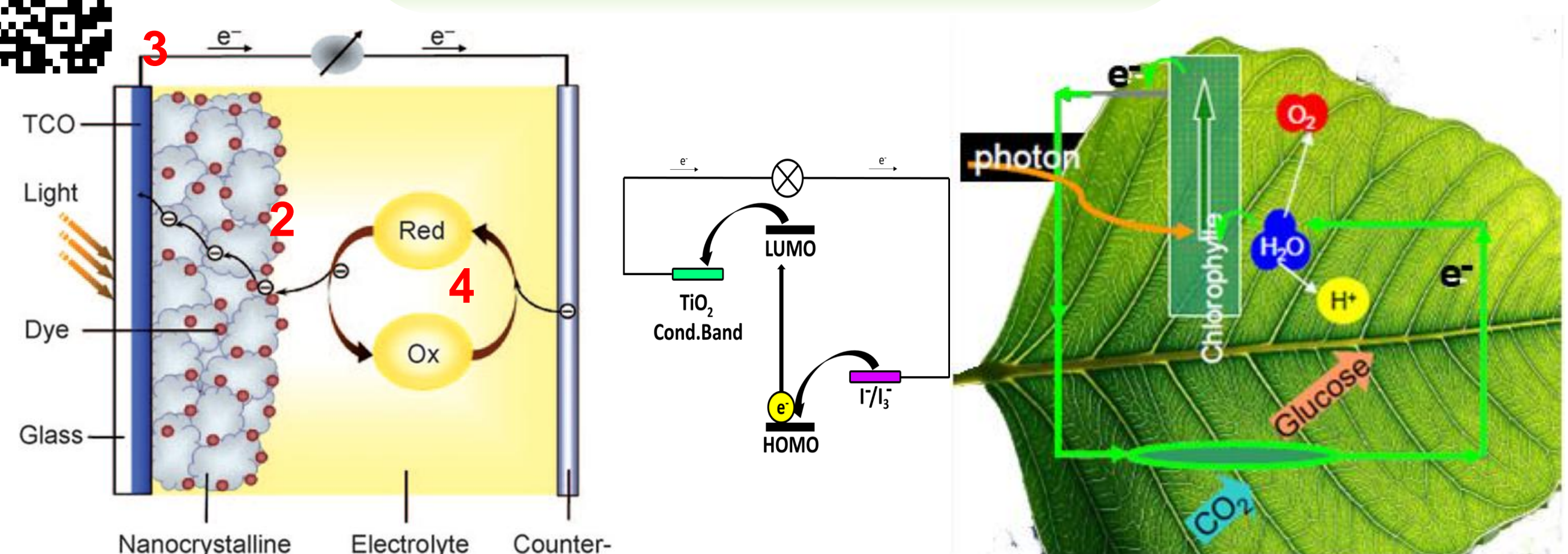
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BACKGROUND: Dye-sensitized solar cells (DSSCs) were invented in 1991 by M. Grätzel. They are inspired by photosynthesis. Compared to silicon-based solar photovoltaic cells, their main advantages are low-cost production, performance in diffuse light and multicolor options. DSSCs are becoming more and more interesting with a large variety of dyes, including vegetable dyes extracted from fruits, algae, flowers and leaves, which are used as light harvesting elements.

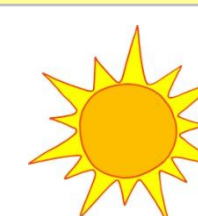
How DSSCs work



1. the dye is the photoactive material of the DSSC
2. the dye absorbs photons from light and uses their energy to promote an electron from the HOMO state to the LUMO state
3. excited electrons are injected into the TiO_2 semiconductor which conducts them into the circuit
4. the electrolyte regenerates the dye

Characterization of dye families

Anthocyanins	β -Carotene	Chlorophylls	Betalains	Organic Dyes	Ruthenium-based dyes
				Examples: -RK1 -GB019 -CEA-INAC Need to be synthesized	
General structure: Blueberries: 				GB019: RK1: 	N719:
extracted vegetable-based dyes				synth. organic	metal complex



Conjugated-bonds and chromophores induce batho- and hyper-chromic shifts. Electrons' delocalization from S(HOMO) to S*(LUMO) state is necessary for injection into TiO_2

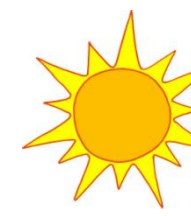
MOTIVATION: Light-weight portable power packs for mobile applications are emerging as one of many industrial applications of DSSCs, with cells made from synthetic dyes and showing over 11% power conversion efficiency in laboratory (1.5% in commercial Gcell module).

We investigated the possibility of using a DSSC incorporating natural pigments extracted from vegetables for charging a cell-phone battery. In order to understand the limitations of such vegetable-based DSSCs, we carried out a series of experiments, including dye extraction and stability measurements, absorption spectroscopy, photovoltaic performance measurements and the charging of a battery. We also used simple models to identify the limitations of our device.

Extraction in ethanol

β -carotene from carrots has no hydroxyl groups, which limits its bonding with ethanol.

Anthocyanin from blueberries on the other hand has many such groups.



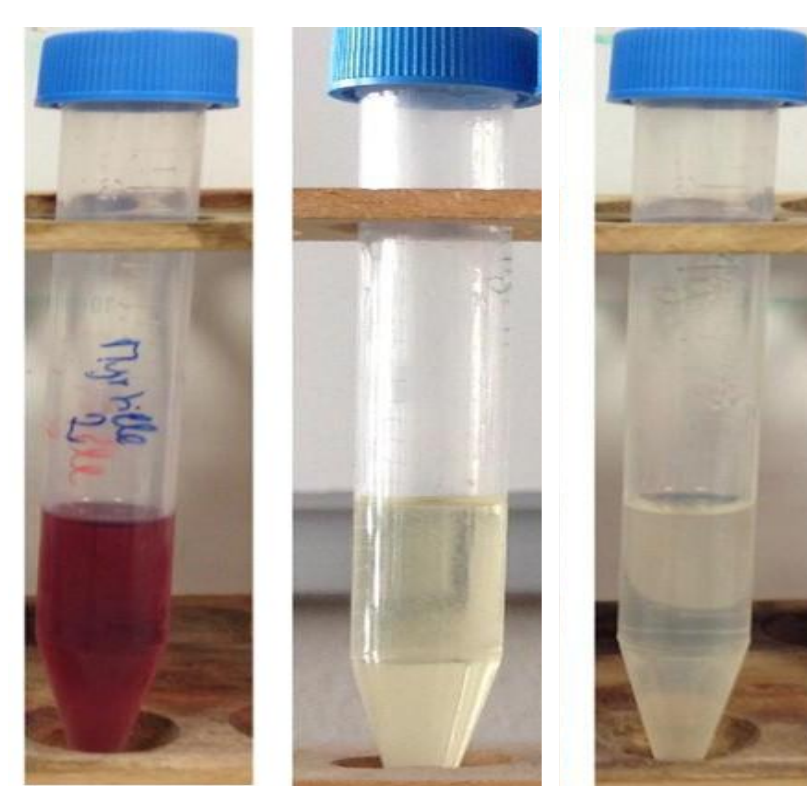
Anthocyanins show:

- + efficient extraction in ethanol
- + harvesting of solar spectrum
- + adsorption to TiO_2 compared to chlorophylls
- +/- dye stability: UV-degradation & precipitation

Choice of vegetable dye

Stability of the dye

Even successfully extracted dyes show degradation over time due to heat and UV-bleaching.



day 0 +3 +20

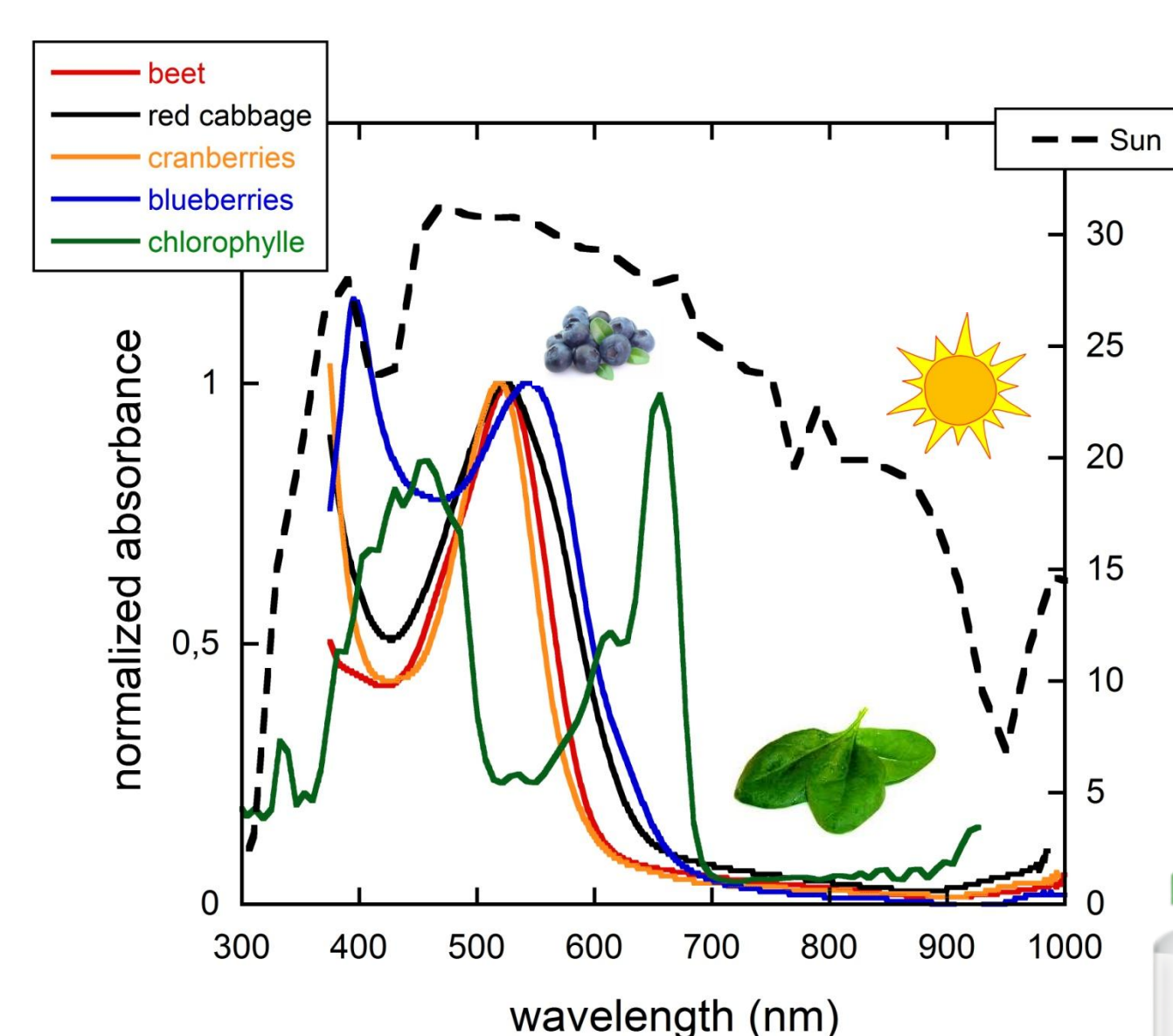
40 ° C + UV-light

Kept in the fridge, the dye does not age as quickly.

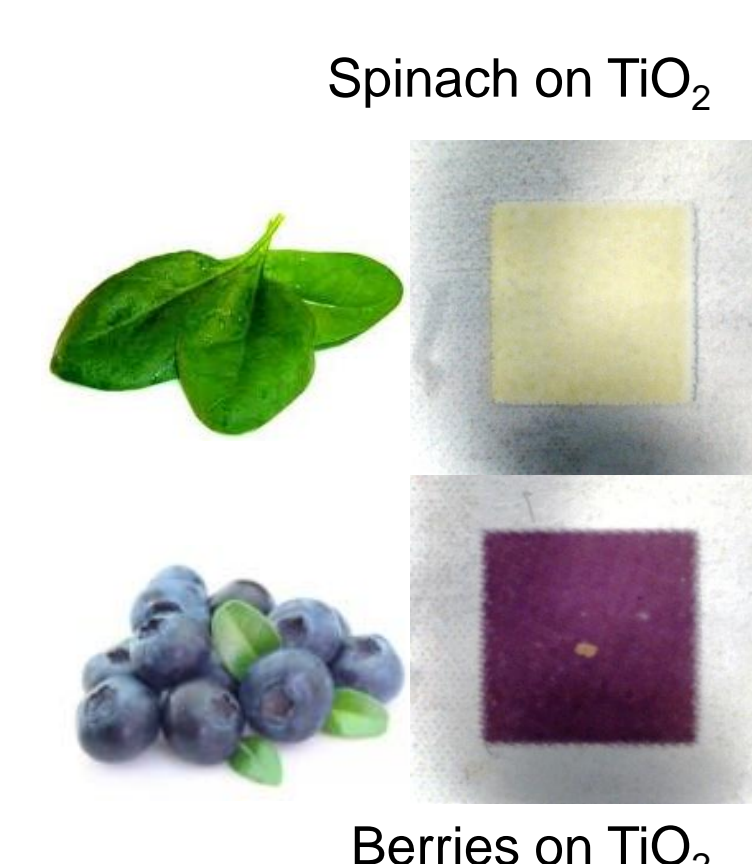


+2 months
4 ° C + dark

Harvesting sunlight spectrum (AM1.5)



Dye adsorption on TiO_2 electrode



Spinach on TiO_2
Berries on TiO_2
Use of blueberry powder allows experiment repetition

DSSC characteristics



Figures of merit

Photovoltaic parameters of the best performing device for each dye obtained under simulated AM 1.5 illumination.

Dye	J_{sc} /mA.cm ⁻²	V_{oc} /mV	FF %	Yield/ %	Ref
N719	12.6	663	75	6.2	This study
N719	17.8	846	75	11.2	(I)
GB019	10.1	655	66	4.3	This study
GB019	1.02	481	72	0.35	This study
GB019	3.14	520	/	0.64	(I)

Yield η : power efficiency

$$\eta_{max} = P_{elec} / P_{sun} = J_{sc} \cdot V_{oc} \cdot FF / \phi_{sun}$$

Solar irradiance $\phi_{sun} = 100 \text{ mW/cm}^2$, AM1.5

Device performance

What is limiting the yield? (kinetics & nanoeffects)

Yield results from IPCE, the Internal Photon to Collected Electron efficiency
 $IPCE(\lambda) = \eta_{abs}(\lambda) \times \eta_{inj} \times \eta_{coll}$

rate equations

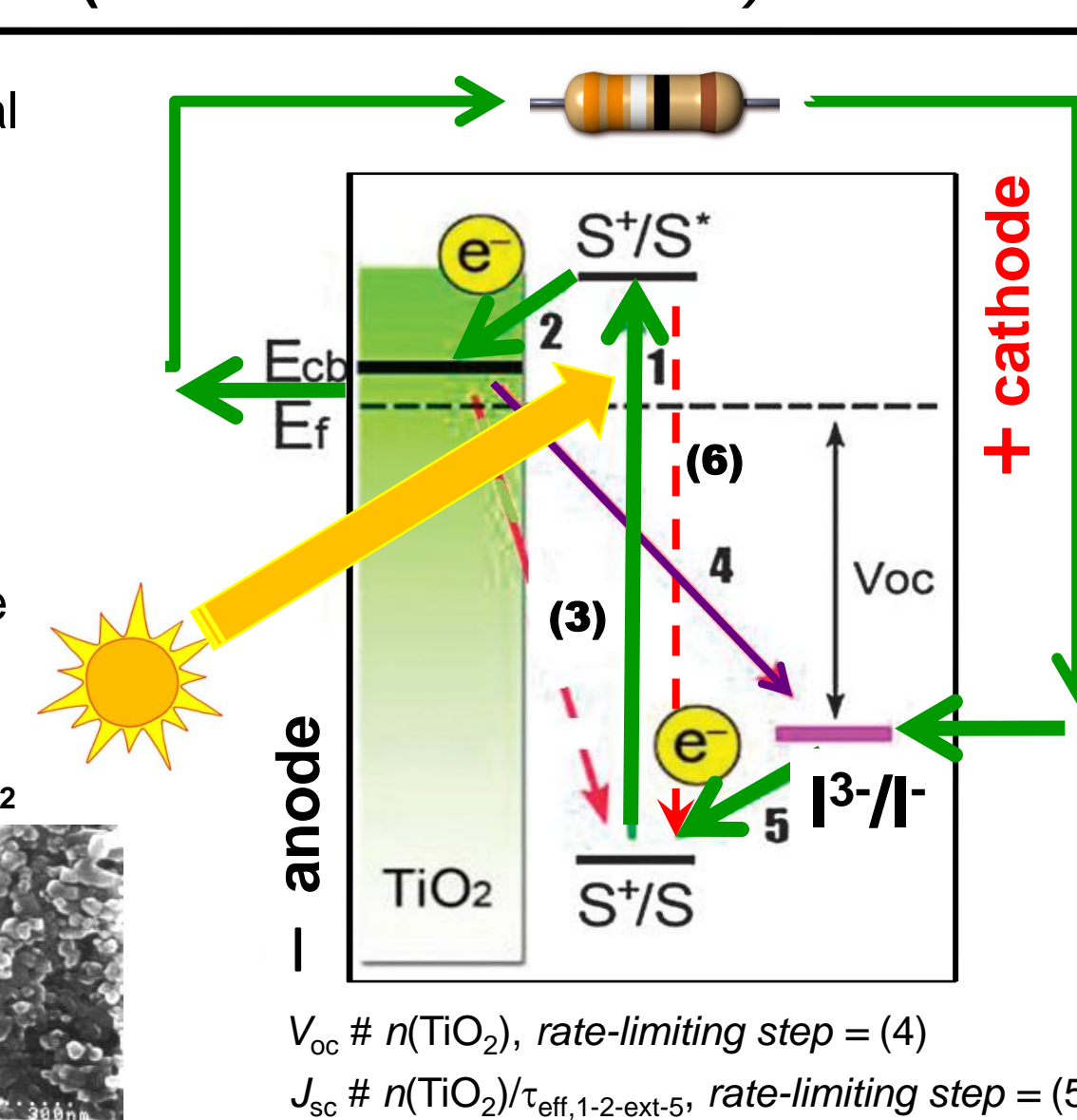
η_{inj} injection efficiency into TiO_2
 η_{coll} collection efficiency into anode
 limited by (4)

fractal nature of nanoporous TiO_2

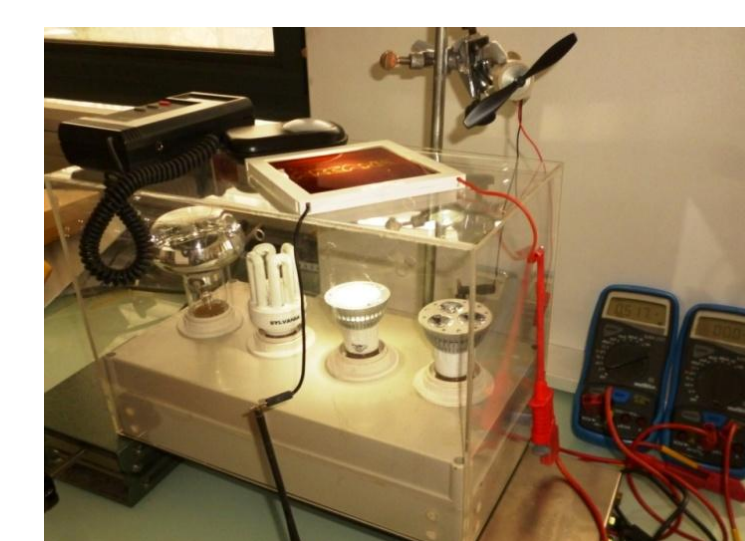
80 m²/g TiO_2 , roughness x 2300

closed pack spheres 20nm ϕ

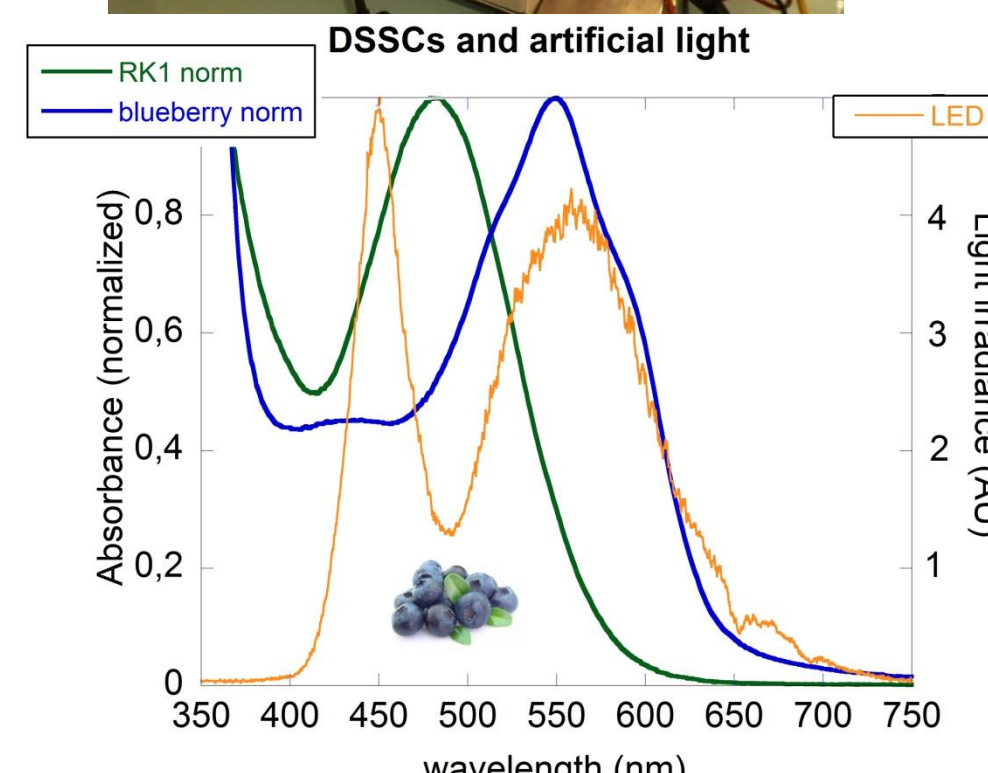
$\eta_{abs}(\lambda)$ absorption efficiency: from 1% to 99%



Performances under diffuse and artificial lights



mW range coupling with 81 cm² demo. cell



Charging a cell phone battery



Charging a 2900 mA.h Li-ion (3.7 V) cell phone battery (iPhone6s) in 2 hours under AM 1.5 (100 mW/cm²) in Grenoble (48° latitude):

- Modules = series association of 5 to 10 cells
- $A(m^2) = \frac{3.7 \cdot Q(mAh)}{\eta(\%) \cdot \Delta t(h)}$

Dye	A (cm ²)	V_{oc} (mV)	N_{serie} cell/module	Yield η (%)	P (mW)	origin	Ref.
N719 (Ru)	0.36	663	1(6)	6.2	2.3	Solaronix®	this study
GCell300 (Ru)	300 (11x27)	723	11		461	G24 Power®	
Arcane (m-free)	81 (1x81)	609	1	0.34	28	Solaronix®	
GB019 (m-free)	0.36	655	1(6)	4.3	1.6	CEA-LEMOH	this study
Blueberries* (vegetable)	0.36	481	1(8)	0.35	0.13	Capsules myrtillus, Fairval®	this study
Blueberries** (vegetable)	520		1(8)	0.64		/	(I)

Blueberry Dye



A0 (this poster)

GB019 A3

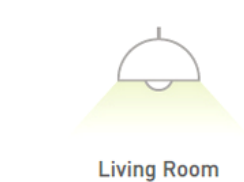
N719 A4

Conclusions



Our study: vegetable-based DSSCs for light-energy conversion and storage

▪ Extrapolation to $\approx 1 \text{ m}^2$ for 5.4 W module necessary for "fast" 2-hour charging of a cell phone: scaling-up issues need to be considered



▪ 100 μW demonstrated with a 0.36 cm² blueberry cell at AM 1.5 (1000 W/m²) standard SUN irradiance



▪ A potential for indoor energy harvesting for applications in the mW regime



▪ Photosynthesis also inspires other "artificial leaves" devices"



this project was endorsed by UNESCO– International Year of Light 2015

